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a textbook of physical chemistry

a textbook for b sc part iii and hons and postgraduate courses of indian universities in this edition i have made major changes in the light of modern concepts introduced in syllabi at the under graduate and postgraduate level as well with matter has also been updated the subject matter has been arranged systematically in a lucid style and simple language new problems and exercises have also been introduced to acquaint the students with trend of questions they expect in the examinations

an advanced level textbook of physical chemistry for the graduate b sc and postgraduate m sc students of indian and foreign universities this book is a part of four volume series entitled a textbook of physical chemistry volume i ii iii iv contents chapter 1 quantum mechanics i postulates of quantum mechanics derivation of schrodinger wave equation max born interpretation of wave functions the heisenberg s uncertainty principle quantum mechanical operators and their commutation relations hermitian operators elementary ideas quantum mechanical operator for linear momentum angular momentum and energy as hermitian operator the average value of the square of hermitian operators commuting operators and uncertainty principle  $x p e t$  schrodinger wave equation for a particle in one dimensional box evaluation of average position average momentum and determination of uncertainty in position and momentum and hence heisenberg s uncertainty principle pictorial representation of the wave equation of a particle in one dimensional box and its influence on the kinetic energy of the particle in each successive quantum level lowest energy of the particle chapter 2 thermodynamics i brief resume of first and second law of thermodynamics entropy changes in reversible and irreversible processes variation of entropy with temperature pressure and volume entropy concept as a measure of unavailable energy and criteria for the spontaneity of reaction free energy enthalpy functions and their significance criteria for spontaneity of a process partial molar quantities free energy volume heat concept gibb s duhem equation chapter 3 chemical dynamics i effect of temperature on reaction rates rate law for opposing reactions of ist

order and second order rate law for consecutive parallel reactions of first order reactions collision theory of reaction rates and its limitations steric factor activated complex theory ionic reactions single and double sphere models influence of solvent and ionic strength the comparison of collision and activated complex theory chapter 4 electrochemistry i ion ion interactions the Debye-Hückel theory of ion-ion interactions potential and excess charge density as a function of distance from the central ion Debye-Hückel reciprocal length ionic cloud and its contribution to the total potential Debye-Hückel limiting law of activity coefficients and its limitations ion size effect on potential ion size parameter and the theoretical mean activity coefficient in the case of ionic clouds with finite sized ions Debye-Hückel-Onsager treatment for aqueous solutions and its limitations Debye-Hückel-Onsager theory for non-aqueous solutions the solvent effect on the mobility at infinite dilution equivalent conductivity  $\lambda$  vs concentration  $c^{1/2}$  as a function of the solvent effect of ion association upon conductivity Debye-Hückel-Bjerrum equation chapter 5 quantum mechanics ii Schrodinger wave equation for a particle in a three-dimensional box the concept of degeneracy among energy levels for a particle in three-dimensional box Schrodinger wave equation for a linear harmonic oscillator its solution by polynomial method zero point energy of a particle possessing harmonic motion and its consequence Schrodinger wave equation for three-dimensional rigid rotator energy of rigid rotator space quantization Schrodinger wave equation for hydrogen atom separation of variables in polar spherical coordinates and its solution principle azimuthal and magnetic quantum numbers and the magnitude of their values probability distribution function radial distribution function shape of atomic orbitals s p d chapter 6 thermodynamics ii Clausius-Clapeyron equation law of mass action and its thermodynamic derivation third law of thermodynamics Nernst heat theorem determination of absolute entropy unattainability of absolute zero and its limitation phase diagram for two completely miscible components systems eutectic systems calculation of eutectic point systems forming solid compounds  $A_xB_y$  with congruent and incongruent melting points phase diagram and thermodynamic treatment of solid solutions chapter 7 chemical dynamics ii chain reactions hydrogen bromine reaction pyrolysis of acetaldehyde decomposition of ethane photochemical reactions hydrogen bromine hydrogen chlorine reactions general treatment of chain reactions ortho-para hydrogen conversion and hydrogen bromine reactions apparent activation energy of chain reactions chain length Rice-Herzfeld mechanism of organic molecules decomposition acetaldehyde branching chain reactions and explosions  $H_2/O_2$  reaction kinetics of one intermediate enzymatic reaction Michaelis-Menten treatment evaluation of Michaelis-Menten constant for enzyme-substrate binding by Lineweaver-Burk plot and Eadie-Hofstee methods competitive and non-competitive inhibition chapter 8 electrochemistry ii ion transport in solutions ionic movement under the influence of an electric field mobility of ions ionic drift velocity and its relation with current density Einstein relation between the absolute mobility and diffusion coefficient the Stokes-Einstein relation the Nernst-Einstein equation Walden's rule the rate process approach to ionic migration the rate process equation for equivalent conductivity total driving force for ionic transport Nernst-Planck flux equation ionic drift and diffusion potential the Onsager phenomenological equations the basic equation for the diffusion Planck-Henderson equation for the diffusion potential

a textbook for B.Sc. classes as per the UGC model syllabus the book is visually beautiful and authors communicate their enthusiasm and enjoyment of the subject in every chapter this textbook is currently in use at hundreds of colleges and universities throughout the country and is a national best seller there are hundreds of computer-generated coloured diagrams graphs photos and tables

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Essentials of Physical Chemistry is a classic textbook on the subject explaining fundamentals concepts with discussions illustrations and exercises with clear explanation

systematic presentation and scientific accuracy the book not only helps the students clear misconceptions about the basic concepts but also enhances students ability to analyse and systematically solve problems this bestseller is primarily designed for b sc students and would equally be useful for the aspirants of medical and engineering entrance examinations

volume five covers dynamics of chemical reactions statistical thermodynamics and macromolecules in five chapters such as adsorption chemical kinetics photo chemistry statistical thermodynamics and macromolecules

this book is a physical chemistry textbook that presents the essentials of physical chemistry as a logical sequence from its most modest beginning to contemporary research topics many books currently on the market focus on the problem sets with a cursory treatment of the conceptual background and theoretical material whereas this book is concerned only with the conceptual development of the subject comprised of 19 chapters the book will address ideal gas laws real gases the thermodynamics of simple systems thermochemistry entropy and the second law the gibbs free energy equilibrium statistical approaches to thermodynamics the phase rule chemical kinetics liquids and solids solution chemistry conductivity electrochemical cells atomic theory wave mechanics of simple systems molecular orbital theory experimental determination of molecular structure and photochemistry and the theory of chemical kinetics

unit i thermodynamics i a recapitulation of thermodynamic terms system surrounding types of system closed open isolated thermodynamic variables intensive extensive properties thermodynamic processes isothermal adiabatic isobaric cyclic reversible irreversible state function path functions properties of state functions exact differential cyclic rule integrating factor concept of heat work 3l b statements of first law of thermodynamics definition of internal energy enthalpy heat capacity at constant volume at constant pressure joule thomson experiment joule thomson coefficient inversion temperature calculations of  $w$   $q$   $\Delta e$   $\Delta h$  for expansion of gases for isothermal adiabatic conditions for reversible process carnot s cycle its efficiency thermodynamic scale of temperature 5l c thermochemistry heat of reaction standard states relation between heat of reaction at constant volume at constant pressure hess s law of constant heat of summation its applications bond dissociation energy its calculations from thermochemical data variation of heat of reaction with temperature kirchoff s equation 4l unit ii thermodynamics ii a second law of thermodynamics need for second law of thermodynamics statements of second law of thermodynamics concept of entropy entropy as a state function of  $v$   $t$   $p$   $t$  entropy change in phase change for ideal gas entropy as criteria of spontaneity equilibrium 4l b free energy functions helmholtz free energy a gibb s free energy  $g$  their properties standard free energies effect of temperature on free energy gibb s helmholtz equation its applications a  $g$  as criteria for thermodynamic equilibrium 4l c system of variable composition partial molar quantities chemical potential van t hoff s reaction isotherm relation between standard free energy change equilibrium constant thermodynamic derivation of law of mass action effect of temperature on equilibrium constant reaction isochore 4l unit iii phase equilibria a phase rule statement of phase rule definition of phase component and degree of freedom derivation of phase rule clapeyron equation its application in deciding slopes of line for two phase equilibria applications of phase rule to two phase equilibria of i water system ii sulphur system iii pb ag system 6l b liquid liquid mixtures ideal liquid mixtures raoults law of ideal solutions henry s law non ideal systems azeotropes hcl h<sub>2</sub>o ethanol water system partial miscible liquids phenol water system trimethylamine water nicotine water system lower upper consolute temperature effect of impurity immiscible liquids steam distillation nernst distribution law limitations deviations applications 6l unit iv solid state laws of crystallography i law of constancy of interfacial angles ii law of rationality of indices iii law of symmetry symmetry of elements in crystals unit cell space lattice orientation of lattice plane miller indices bravais lattices crystal systems x ray diffraction by crystal derivation of braggs equation

this comprehensive textbook now in its second edition is mainly written as per the latest syllabi of physical chemistry of all the leading universities of india as well as the new syllabus recommended by the ugc this thoroughly revised and updated edition covers the principal areas of physical chemistry such as thermodynamics quantum chemistry molecular spectroscopy chemical kinetics electrochemistry and nanotechnology in a methodical and accessible style the book discusses classical irreversible and statistical thermodynamics and statistical mechanics and describes macroscopic chemical systems steady states and thermodynamics at a molecular level it elaborates the underlying

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calculations approach strong mathematical rigor has been applied and a complementary physical treatment given to make students strong in the applied aspects of thermodynamics problem solving presentation 195 solved examples and 269 unsolved problems have been given hints to difficult problems have been give too concept checking review questions have been given at the end of every chapter coverage on thermodynamic discussion of eutectics solid solutions and phase separation

aspects of the learning process are fully supported including the understanding of terminology notation mathematical concepts and the application of physical chemistry to other branches of science building on the heritage of the world renowned atkins physical chemistry quanta matter and change gives a refreshing new insight into the familiar by illuminating physical chemistry from a new direction book jacket

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